

WS 4.13 Hydrolysis of Salts

Name: _____

Write the net ionic equation for the hydrolysis each of the following salts (if any). State if the solution will be acidic, basic or neutral.

| | |
|-----|------------------------------------|
| 1. | NaCH_3COO |
| 2. | NH_4Cl |
| 3. | $(\text{NH}_4)_2\text{SO}_4$ |
| 4. | $\text{Ca}(\text{NO}_3)_2$ |
| 5. | $\text{Fe}_2(\text{SO}_4)_3$ |
| 6. | MgO |
| 7. | SO_3 |
| 8. | KBr |
| 9. | CrCl_3 |
| 10. | $\text{NH}_4\text{CH}_3\text{COO}$ |
| 11. | NH_3 |
| 12. | K_2HPO_4 |
| 13. | KH_2PO_4 |
| 14. | KHSO_4 |
| 15. | H_2O_2 |

