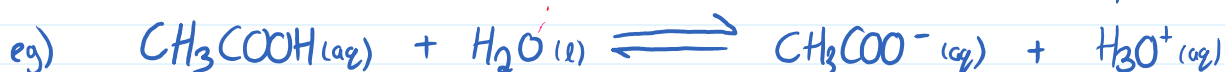


4.8 Ka & Kb

Ka ∴ Acid Ionization Constant

- for weak acids (Remember they don't ionize 100%)

- Key expression for weak acids (Not top 6 in + ble)



- stronger acid = greater Ka (greater ionization, ∴ more $[\text{H}_3\text{O}^+]$ on top)

Kb ∴ Base Ionization Constant



- stronger base = greater Kb!

Example calculation: What is the $[\text{H}_3\text{O}^+]$ of 0.25M CH_3COOH ?



I

C

E

Do 4.8 Pg. 128 #31-34

